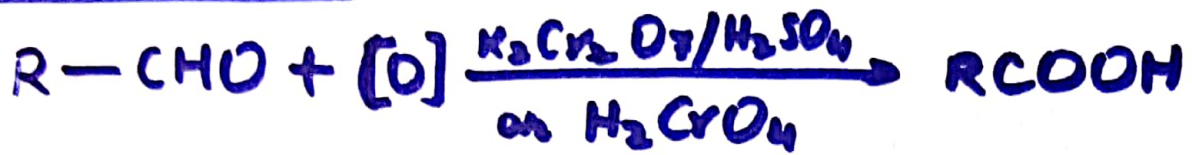
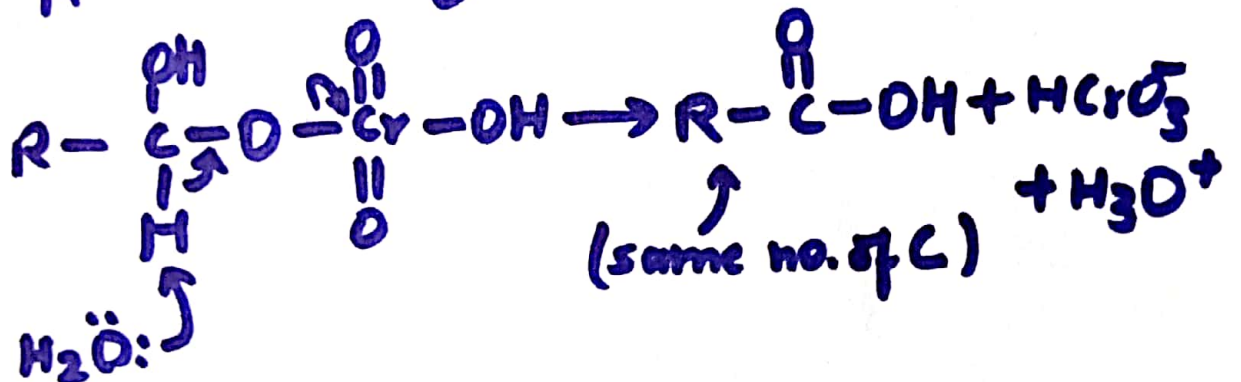
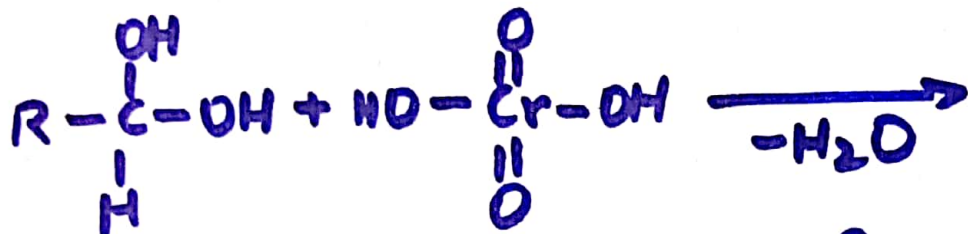


ALDEHYDES & KETONES

OXIDATION OF ALDEHYDES



Ald. first gets hydrated to a gem-diol.



easily oxidised by mild oxidising agents
So they are strong reducing agents.

Readily reduce (i) Tollens' reagent to Ag
(ii) Fehling's solution or Benedict soln
to red ppt. of Cu_2O

Aromatic aldehydes do not reduce
Fehling's soln.

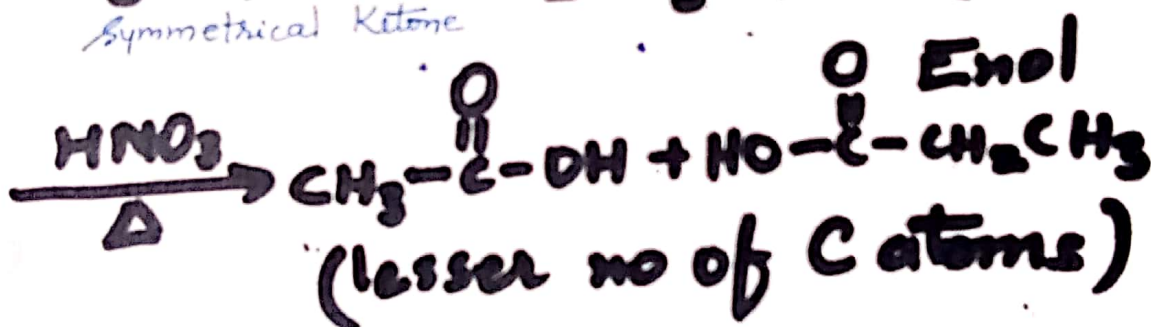
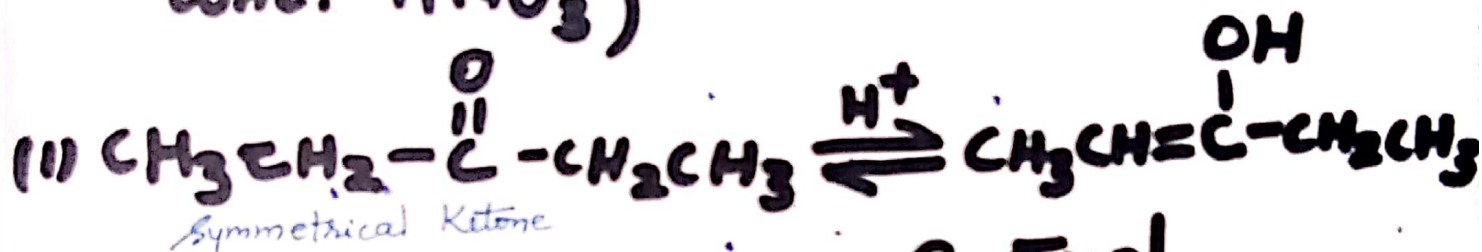
OXIDATION OF KETONES

C-C cleavage required so ketones are resistant to oxdn.

oxidised to carboxylic acid

under drastic conditions →

(acidified $K_2Cr_2O_7$, alk. $KMnO_4$, conc. HNO_3)



(ii)

